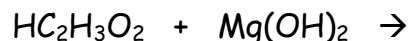


1. Give the word equation for the neutralization reaction of an acid and a base.
2. Complete these equations:



3. A \_\_\_\_\_ is a laboratory method used to determine the concentration of an acid or a \_\_\_\_\_ in solution by performing a \_\_\_\_\_ reaction with a standard solution.
4. At the \_\_\_\_\_ of the titration, the indicator changes color, which indicates neutralization. Once neutralized, moles of \_\_\_\_\_ and moles of \_\_\_\_\_ are equal.
5. In a titration of HCl with NaOH, 100.0 mL of the base was required to neutralize 20.0 mL of 5.0 M HCl. What is the molarity of the NaOH? (Be sure to write the neutralization reaction.)
  
6. In a titration of  $\text{H}_2\text{SO}_4$  with NaOH, 60.0 mL of 0.020 M NaOH was needed to neutralize 15.0 mL of  $\text{H}_2\text{SO}_4$ . What is the molarity of the acid? (Be sure to write the neutralization reaction.)
  
7. If 10.0 mL of 0.300 M KOH are required to neutralize 30.0 mL of gastric juice (HCl), what is the molarity of the gastric juice?